

Combined TD-DFT-SOS-CIS(D) Study of BOPHY Derivatives with Potential Application in Biosensing

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7 A Combined TD-DFT-SOS-CIS(D) Study of
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11 BOPHY Derivatives with Potential Application in
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47 ABSTRACT
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50 A set of thirteen BOPHY dyes is studied through a hybrid TD-DFT-SOS-CIS(D) approach
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52 accounting for solvent effects, in order to shed light onto the structure-property relationships of
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54 these recently developed chromophores. In a first step, we calculate the absorption fluorescence
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56 crossing points with refined TD-DFT models considering the influences of both vibrational and
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3 solvent contributions. We found that the systematic overestimation of the 0-0 energies is
4 effectively reduced by combining PCM-TD-DFT with a scaled opposite spin-configuration
5 interaction singles with a double correction [SOS-CIS(D)]. Next, for a representative system, the
6 vibrationaly resolved spectrum within the harmonic approximation is computed on the basis of
7 TD-DFT vibrational signatures and an excellent match with experiment is found. Finally, the
8 influence of different lateral groups on the spectroscopic properties is rationalized by
9 investigating charge transfer parameters and examining electronic density difference maps. It is
10 found that one can tune the position of the absorption/emission maxima by a judicious choice of
11 the lateral substituents or by using π -extended segments. The largest absorption and emission
12 wavelengths as well as the largest Stokes shifts are obtained for BOPHYs containing strong
13 electron-donor dimethylaminophenyl groups attached to the α -positions of the pyrrole units
14 through vinyl linkers, making these chromophores promising candidates for bioluminescence
15 applications.
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3 INTRODUCTION
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7 During the past decades, technological advances in the field of bioimaging have allowed
8 monitoring several intracellular processes that were previously invisible to human eyes¹⁻³. In this
9 framework, the search for novel molecules exhibiting appropriate optical signatures for
10 fluorescent probe applications, i.e., high fluorescent quantum yield, photo and chemical stability,
11 biocompatibility with the medium, and large Stokes shift, has been particularly intense⁴⁻⁷. The
12 strong UV-absorption and high fluorescence quantum-yield of BODIPY (4,4-difluoro-4-bora-
13 3a,4a-diaza-s-indocene) dyes have prompted several research groups to synthesize structurally
14 similar compounds that may present enhanced luminescent properties^{8,9}. Among these new
15 fluorophores, bis(difluoroboron)-1,2-bis((pyrrol-2-yl)methylene)hydrazine (BOPHY, see
16 Scheme 1) derivatives, that contain two borondifluoride units bridged by a coupled pair of Schiff
17 bases, arise as promising candidates as they exhibit large extinction coefficients¹⁰, emission
18 quantum yields close to unity¹⁰ and deep-red fluorescence emission¹¹⁻¹³. Since their discovery,
19 several structural modifications of BOPHYs have been conducted in order to boost their
20 emissive properties, including the incorporation of halogens¹⁴, perylene¹⁴, and ferrocene units¹⁵
21 to the bis(difluoroboron) scaffold. In addition, a (2-hydroxyphenyl)ethylidene-substituted
22 BOPHY probe for Cu²⁺ has been recently synthesized by Li and coworkers¹⁶, which is able of
23 first coordinating the target, and next, through the hydrolysis of the C=N bond, generating a
24 formyl-substituted fluorescent BOPHY.
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50 To the best of our knowledge, there has been only one previous theoretical study
51 concerning BOPHYs¹⁷. This work concerned the simulation of one- and two-photon absorption
52 bands of bis(BF₂) complexes containing naphthyridine derivatives and it revealed the good
53 performance of the State Specific Polarizable Continuum Model (SS-PCM) approach in the
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3 estimation of the inhomogeneous broadening of the spectra. This initial study represents a
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5 valuable foundation for *ab initio* calculations of the optical properties involving BOPHY
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7 molecules in solution. In the present contribution, we intend to study the optical properties of an
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9 extended set of BOPHY derivatives, using a Time-Dependent Density Functional Theory (TD-
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11 DFT) protocol combined with a scaled opposite spin-configuration interaction singles with a
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13 double correction [SOS-CIS(D)] approach. In the first stage, we focus on the 0-0 energies (E^{0-0})
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15 calculations, given the interest of this parameter that allows physically well-grounded
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17 theory/experiment comparisons. E^{0-0} can be defined as the difference of the excited-state (ES)
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19 and ground-state (GS) energies taken at their respective optimal geometries and corrected for the
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21 difference of zero-point vibrational effects between the two states. E^{0-0} can be accessed
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23 computationally and compared to the experimental absorption and fluorescence crossing point
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25 (AFCP, see Computational Details)¹⁸. Second, we simulate the vibrationally resolved spectra
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27 within the harmonic approximation on the basis of vibrational signatures. In this context, we note
28
29 that vibronic TD-DFT calculations have been previously successfully employed to obtain the
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31 band shapes of (aza-) BODIPYs¹⁹, BODIPY-cyanines²⁰ and Ladder-type- π -conjugated dyes²¹.
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33 Finally, we rationalize the effect of different lateral groups in the absorption/emission properties
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35 of the studied BOPHYs. To this end, we study electronic density difference (EDD) plots and
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37 charge transfer (CT) parameters. This work aims at providing some valuable guidelines to
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39 experimental teams working on the synthesis of BOPHY structures with improved optical
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41 properties.
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COMPUTATIONAL DETAILS

All DFT/TD-DFT calculations were carried out with the Gaussian09 software²², tightening self-consistent field convergence thresholds (10^{-10} a.u.). For all these calculations the M06-2X²³ hybrid exchange-correlation functional has been used. This choice of the exchange-correlation functional is dictated by our previous works on (aza-) BODIPYs¹⁹, BODIPY-cyanines²⁰ and Ladder-type- π -conjugated dyes²¹, all boron containing systems. Indeed, while this functional provides blue-shifted transition energies compared to experiment -an error common to all functionals when used in an adiabatic TD-DFT context for cyanine-like derivatives²⁴- the obtained structural and vibrational parameters were found accurate. All GS and ES geometry optimizations were followed by vibrational frequency calculations, where the absence of imaginary frequencies demonstrates the stability of the minimum of the structures. Two atomic basis sets, namely, 6-31G(d) for geometry optimizations and frequency calculations, and 6-311+G(2d, p) for electronic transitions have been applied. In order to reproduce the solvation effects, the Polarizable Continuum Model (PCM) is employed, using dichloromethane as solvent as in the experiments. In the PCM model, the solvent is described as a polarizable continuum medium characterized by its dielectric permittivity ϵ and other macroscopic parameters. The interaction with the solute is modeled through a charge distribution on the cavity^{25,26}. The *equilibrium* limit is considered for geometry optimizations, whereas the *non-equilibrium* limit is applied for calculations of both absorption and emission energies. In the framework of excited-state calculations with PCM²⁷, there are several approaches to account for the density reorganization, either using the transition matrix in the so-called Linear Response (LR) approach^{28,29}, or the one-particle density matrix using the corrected Linear Response (cLR)³⁰ or State-Specific (SS)³¹ schemes. A combined protocol including TD-DFT and SOS-CIS(D) has

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been used to reduce the overestimation obtained in the TD-DFT excitation energies. Grimme and coworkers evaluated the performance of CIS(D) for the prediction of 0-0 excitation energies of large organic molecules, and they concluded that it represents a reliable and efficient approach given its size-consistency and relatively low computational cost³²⁻³⁴. However, it is clear that the range of CIS(D) application is limited by some requirements including a good description of the ground- and excited-states by the HF determinant and the CIS reference function, respectively^{35,36}. In addition, it was reported that CIS(D) sometimes yields significant blue shifts of more than several tenths of an eV, due to near-degeneracy effects³⁷. Moreover, the CIS(D) method and its variations are in practice much less efficient than TD-DFT for determining ES vibrational structures and geometries, and for this reason the CIS(D) calculations are often coupled to a TD-DFT protocol¹⁹ in which the former is used to determine transition energies and the latter the structural parameters. The SOS-CIS(D) variation used herein involves a single scaling parameter that influences the opposite-spin direct term of CIS(D) starting from the one-parameter scaled opposite spin MP2 ground-state³⁸. This method has been shown to yield very satisfying results for several classes of dyes^{19, 20, 21, 39-41}. The SOS-CIS(D) calculations have been computed with the Q-Chem package⁴² with the 6-311+G (2d,p) basis set and by applying the resolution of the identity scheme, with a triple- ζ auxiliary basis set. Both CIS(D) and SOS-CIS(D) partially account for the impact double excitations that are known to be important in boron difluoride dyes²⁴. The AFCP energies were calculated with a recently designed protocol⁴³⁻⁴⁵, that is briefly summarized below.

In the gas phase, a (idealized) vertical absorption ($E^{\text{vert-abs}}$) can simply be defined as the difference between the ES and GS energies (E^{ES} and E^{GS} , respectively) at the optimal ground-state geometry (R^{GS}),

$$E^{\text{vert-abs}} = E^{\text{ES}}(R^{\text{GS}}) - E^{\text{GS}}(R^{\text{GS}}) \quad (1)$$

whereas the vertical fluorescence ($E^{\text{vert-fluo}}$) can be calculated in a similar manner, but considering the excited-state geometry.

$$E^{\text{vert-fluo}} = E^{\text{ES}}(R^{\text{ES}}) - E^{\text{GS}}(R^{\text{ES}}) \quad (2)$$

The adiabatic energy (E^{adia}) is the difference of total electronic energies computed for the ES and GS in their corresponding optimal geometries:

$$E^{\text{adia}} = E^{\text{ES}}(R^{\text{ES}}) - E^{\text{GS}}(R^{\text{GS}}) \quad (3)$$

It can be calculated alternatively by combining vertical transition energies with the geometrical reorganization energies (E^{reorg}), that measure the energy difference determined for a given single electronic state when considering the optimal GS and ES geometries:

$$E^{\text{adia}} = \frac{1}{2}[E^{\text{vert-fluo}} + E^{\text{vert-abs}}] + \frac{1}{2}[E^{\text{reorg-GS}} - E^{\text{reorg-ES}}] \quad (4)$$

In this latter equation the first term is dominant and we can make the following approximation:

$$E^{\text{adia}} \simeq \frac{1}{2}[E^{\text{vert-fluo}} + E^{\text{vert-abs}}] \quad (5)$$

At the right-hand side of this last equation, we can observe the average between the vertical absorption and emission. Next, we determine the difference of zero-point vibrational energy (E^{ZPVE}) between the ES and GS,

$$\Delta E^{\text{ZPVE}} = E^{\text{ZPVE}}(R^{\text{ES}}) - E^{\text{ZPVE}}(R^{\text{GS}}) \quad (6)$$

to reach the 0-0 energy (E^{0-0}) we add E^{adia} and ΔE^{ZPVE} ,

$$E^{0-0} = E^{\text{adia}} + \Delta E^{\text{ZPVE}} \quad (7)$$

E^{0-0} can be compared to the experimentally measured absorption-fluorescence crossing point (AFCP), and it therefore offers a solid basis for theory-experiment comparisons.

For a representative system, the vibrationally resolved spectrum within the harmonic approximation was computed by using the FCclasses program on the basis of TD-DFT vibrational signatures⁴⁶⁻⁴⁹. The reported spectrum was simulated by using a convoluting Gaussian function presenting a half width at half-maximum (HWHM) that was adjusted to allow direct comparison with the experimental spectrum (HWHM was set to 0.085 eV and 0.055 eV for absorption and fluorescence, respectively). A maximal number of 25 overtones for each mode and 20 combination bands for each pair of modes were included in the calculation. The maximum number of integrals to be computed for each class was set to 10^6 , which provided optimal spectra.

To visualize the electronic density redistribution and the possible intramolecular charge transfer we represented the difference between the total densities of the excited- and ground-states (EDD) as provided by TD-DFT,

$$\Delta\rho(r) = \rho^{\text{ES}}(r) - \rho^{\text{GS}}(r) \quad (8)$$

All EDD maps and HOMO-LUMO isosurfaces have been plotted with the Chemcraft code⁵⁰, considering a contour threshold of 0.001 a.u. for the former and 0.03 a.u. for the latter. Additionally, the amount of CT (q_{CT} in |e|), CT distance (d_{CT} in Å), and CT dipole (μ_{CT} in Debye) have been calculated according to the methodology proposed by Le Bahers and coworkers^{51,52}.

RESULTS AND DISCUSSION

Statistical Analysis

In previous works dealing with the modeling of spectroscopic properties of BODIPY derivatives, the PCM-TD-DFT protocol enabled to obtain theoretical results in good qualitative agreement with experimental data, but generally too large on the absolute energy scale¹⁸⁻²¹. Here, we have tested two models for transition energies, TD-DFT and SOS-CIS(D), in both gas phase or combined with three solvent approaches (LR, cLR, and SS) in order to check which approach is the most effective in reproducing the experimental AFCP values for the thirteen BOPHY dyes represented in Scheme 2. The experimental AFCP energies and their theoretical counterparts are listed in Table 1, whereas the results of a statistical analysis providing the mean absolute error (MAE), mean square error (MSE), standard deviation (SD), maximal positive and negative deviations [Max(+) and Max(-)], as well as theory/experiment linear determination coefficients (R^2) can be found in Table 2. First of all, let us compare the experimental AFCP values, and the 0-0 energies obtained by TD-DFT and SOS-CIS(D) in gas phase (Figure 1). The obtained results evidence that without considering solvent effects the AFCP trend is rather well reproduced by TD-DFT with a R^2 of 0.935 but at a cost of a very large MAE (0.426 eV) and a sizeable MSE (0.185 eV). Conversely, the gas-phase SOS-CIS(D) calculations provide much smaller average errors (MAE of 0.094 eV and MSE of 0.013 eV), but do not enable to accurately reproduce the observed trend ($R^2 = 0.869$). This is mainly due to an overestimation of the 0-0 values for those BOPHYs that absorb light at small energies (**5**, **6** and **7**). Similar deviations in the SOS-CIS(D) results have been evidenced by Goerigk and Grimme for a panel of twelve chromophores of different nature³².

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Next, we carried out the same study by TD-DFT methodologies incorporating the solvent effects via the state specific (SS-TD-DFT), linear response (LR-TD-DFT) and corrected linear response (cLR-TD-DFT) schemes (Figure 2). An improvement in terms of correlation is observed. The cLR-TD-DFT approach leads us to the highest R^2 value of 0.968, whereas SS and LR provide values of 0.965 and 0.955, respectively. As in the gas phase, the incorporation of SOS-CIS(D) deteriorates these coefficients, but only slightly here, with values of 0.941 for LR, 0.950 for cLR and 0.936 for SS. Again, the MAE and MSE values clearly show the positive impact of the inclusion of the SOS-CIS(D) corrections in terms of accuracy. Indeed, compared to TD-DFT the MAE (MSE) values decrease from 0.221 eV (0.052 eV) to 0.199 eV (0.046 eV) for LR, from 0.431 eV (0.188) to 0.061 eV (0.005 eV) with cLR, and from 0.521 eV (0.277 eV) to 0.104 eV (0.014 eV) with SS. In Figure 2, one clearly sees that TD-DFT overestimates the experimental AFCP energies, irrespective of the selected solvent models, which was, as stated above, expected from previous simulations of fluoroborates^{19,21}. SS-SOS-CIS(D) also slightly overshoots the experimental data, whereas we found a significant underestimation with LR-SOS-CIS(D). Overall, the most accurate approach is cLR-SOS-CIS(D) that delivers only a trifling average underestimation compared to the experimental results. It is also interesting to note that to obtain both accurate absolute values and high correlation, one needs the inclusion of both solvent effects and the impact of "double excitation" in the calculation, highlighting the challenge of theoretically modeling BOPHY with high accuracy.

Analysis of the band topology

Globally, the shapes of the absorption and emission bands of BOPHYs are rather independent of their substitution patterns, suggesting that vibronic couplings can be attributed to the BOPHY core rather than the auxochromes. Consequently, we investigated the vibronic couplings of molecule **1** as a representative example, by applying the FC approximation, and the spectra are presented in Figure 3. These results are perfectly converged with FC factors exceeding 0.95 for both absorption and emission. It is rather obvious that theory reproduces very nicely both the position and the band shapes compared to experiment of Ref. 11, the only slight difference being the relative heights of the two maxima in the absorption spectrum that are inverted by the model. Unlike the BODIPY dyes, where the emission profile is the mirror image of the absorption one, a clear difference exists between them for BOPHY, and this is well reproduced by our calculations. Interestingly the computed absolute molar extinction coefficient, ϵ , that is $4.1 \cdot 10^4 \text{ M}^{-1} \text{ cm}^{-1}$, has the correct order of magnitude compared to the experiment, $7.0 \cdot 10^4 \text{ M}^{-1} \text{ cm}^{-1}$, which also corroborates the quality of the FC model in the present case. Additionally, we have analyzed the transitions playing a key role in the band shapes. For the absorption spectrum, one mode is highly relevant: an excited-state vibration at 1533 cm^{-1} , that is responsible for both, the second maximum (one quantum) and the significant broadening at higher energies (two quanta). This mode is mainly related to the intense stretching of the central N-N bond. This explains why the typical band shape of BODIPY is not recovered in BOPHY: the new N-N linkage plays a key role in the vibronic couplings, and hence, in the band shape. For the emission, the marked shoulder is principally due to a ground-state vibration at 1397 cm^{-1} (see Supporting Information). This mode involves the central N-N bond as well, but also includes a marked wagging of the hydrogen atoms. The different natures of the coupling modes

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3 for absorption and emission are consistent with the unsymmetrical band topologies measured
4 experimentally.
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10 11 12 *Structure-Property Relationships* 13

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16 The BOPHY framework is characterized by two fused BF₂-containing six-membered
17 rings in the center and two five-membered pyrrole moieties at the periphery. Whereas the pyrrole
18 and BF₂-containing units are coplanar, there exists a central dihedral angle of ca. 28.5° in all
19 chromophores irrespective of their C₂ (**1**, **2**, **2-Ph**₂, **2-tBu**₂, **2-OMe**₂, **3**, **4**, **7**) or C₁ (**2-Ph**, **2-tBu**,
20 **2-OMe**, **5**, **6**) symmetry (see Table 3). It seems that this dihedral angle remains unaltered despite
21 the addition of different substituents, in contrast with the BODIPY fluorophores in which the
22 incorporation of lateral substituents usually favors structural planarity⁵³. In all cases, the central
23 dihedral angle is reduced in the excited-state, denoting a greater π-conjugation in S₁ (see Table
24 3). Additionally, a reduction of the central N-N bond distance occurs compared to the ground-
25 state, whereas the B-N distances do not present significant variations.
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40 The HOMO-LUMO plots in Figure 4 for six representative BOPHY systems reveal that
41 the HOMO is mainly spread over the nitrogen atoms and peripheral rings, with the exception of
42 compounds **5** and **7** in which dimethylaminophenyl and methoxyphenyl groups connected
43 through vinyl linkers also participate significantly. In contrast, the LUMO is distributed over the
44 pyrrole units and the *meso*- carbon atoms, i.e., those adjacent to the central N-N bond. Recently,
45 Nemykin and coworkers attempted to introduce a cyano substituent to this *meso*- position, but
46 the dramatic reduction of the nitrogen atom basicity resulted in a partial extrusion of the BOPHY
47 core⁵⁴.
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3 The incorporation of methyl groups in the α - and β -positions of the pyrrole units (Scheme
4 1) of **1** generates the tetra-methyl substituted compound **2** and causes a redshift of +20 nm (-
5 0.132 eV) in absorption, and +20 nm (-0.110 eV) in emission. The subsequent addition of ethyl
6 groups to the β -positions leading to **3** induces additional bathochromic shifts of +34 nm (-0.199
7 eV) and +12 nm (-0.062 eV) in the absorption and emission, respectively. Similar red-shift
8 magnitudes are obtained when alkyl groups are added to the β -positions of BODIPYs⁸.
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18 The inclusion of a phenylacetylene group in **2** leads to compound **2-Ph** which yields
19 absorption and emission redshifts of +12 nm (-0.073 eV) and +24 nm (-0.121 eV), respectively,
20 evidencing the effective but rather limited electronic communication between the BOPHY
21 moiety and the phenyl group through the ethynyl linker. However, the electronic density
22 difference (EDD) maps (Figure 5) reveal that this linker could actually act as a secondary
23 electron donor group, which is consistent with the absorption [+9nm (-0.055 eV)] and emission
24 [+14nm (-0.072 eV)] red-shifts of similar magnitude observed in **2-*t*Bu**. In this context, let us
25 recall that incorporating *tert*-butyl substituents advantageously limits π - π aggregation in solid-
26 state, which allows to avoid the so-called solid-state emission quenching^{55,56}.
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41 The electron-donor character of the methoxyphenyl group is capable to generate red-
42 shifts in aza-boron-dipyrromethene dyes, when this group is incorporated into α -positions⁵⁷.
43 Analogously, in BOPHY dyes the influence of one methoxyphenyl attached to the pyrrole unit
44 through an ethynyl linker (**2-OMe**) causes a bathochromic shift of +35 nm (-0.204 eV) for the
45 absorption and a +38 nm (-0.186 eV) shift for the fluorescence, with respect to **2**, more
46 pronounced than those observed in **2-Ph** and **2-*t*Bu**. The large q_{CT} value calculated for the **2-
47 OMe** system (0.511 |e|) evidences a synergic effect caused by the ethynyl linker and the
48 methoxyphenyl moiety.
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In 3-dimethylaminophenyl-8-CF₃-BODIPY, the addition of the strong electron-donor amino groups to the pyrrole units triggers an intramolecular charge transfer leading to bathochromic red-shifts greater than +100 nm for both absorption (-0.230 eV) and emission maxima (-0.209 eV)⁵⁸. In the BOPHY case, compounds **5** and **6** -where the electron-donor substituent 4-dimethylaminophenyl is linked through a vinyl connector- present larger bathochromic displacements in comparison to their **2** and **3** analogues. Hence, **5** exhibits an absorption red-shift of +113 nm (-0.567 eV) and an emission red-shift of +182 nm (-0.698 eV) with respect to **2**, whereas **6** is red-shifted in +79 nm (-0.368 eV) in absorption and +173 nm (-0.644 eV) in emission, with respect to **3**. Moreover, these two chromophores display the largest absorption and emission wavelengths of the panel, which can be explained by their pronounced donor-acceptor character, evidenced by high charge transfer dipole (μ_{CT}) values of 10.3 D (**5**) and 10.0 D (**6**) (see Table 4). In addition, the greatest q_{CT} values of our set of dyes are obtained with these unsymmetrical chromophores (0.60 |e| for both **5** and **6**). The EDD plots also reveal the strong variation experienced by the electronic densities of the donor lateral arms, demonstrating the efficient π -conjugation achieved by the vinyl linkers. In contrast, the presence of two methoxyphenyl groups in the α -positions (**7**) decreases the dipole CT character of the molecule, resulting in a smaller emission maximum (611 nm). While the μ_{CT} is not significant for the symmetric dye **7**, we notice that its q_{CT} (0.516 |e|) is significantly smaller than those corresponding to **5** and **6**. Another advantageous feature exhibited by **5** and **6** is their large Stokes shift of 2961 cm⁻¹ and 3028 cm⁻¹, respectively, probably preventing re-absorption of emitted photons that might lead to undesired background interferences⁵⁹. At this point we can state that the very large absorption and emission maxima found for **5** and **6** chromophores, near

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3 the so-called first biological window⁶⁰ and their large Stokes shifts, make them an interesting
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5 starting point for the design of novel bioluminescence agents.
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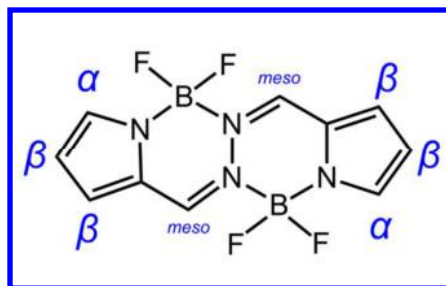
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9 Increasing π -electron conjugation of similar emissive organic chromophores normally
10 results in enhanced fluorescent emission, as well as redshifts in excitation and emission
11 wavelengths⁶¹⁻⁶⁴. This strategy is applied in system **4**, a naphthyridin derivative⁶⁵ where the π -
12 conjugation is extended by fusing the pyrrole units with other aromatic rings. We found that
13 absorption maximum is red-shifted in +54 nm (-0.303 eV), and emission in +95 nm (-0.419 eV),
14 with respect to **2**, as a result of the donor effect of the added rings (see EDD map). In the
15 BODIPY case, by extending the central core, red-shifts of +39 nm (-0.125 eV) and +42 nm (-
16 0.130 eV) are found for absorption and emission, respectively⁶³, while in the case of aza-
17 BODIPYs, smaller displacements of ca. +16 (-0.045 eV) are found for both absorption and
18 emission⁶⁴.
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CONCLUDING REMARKS

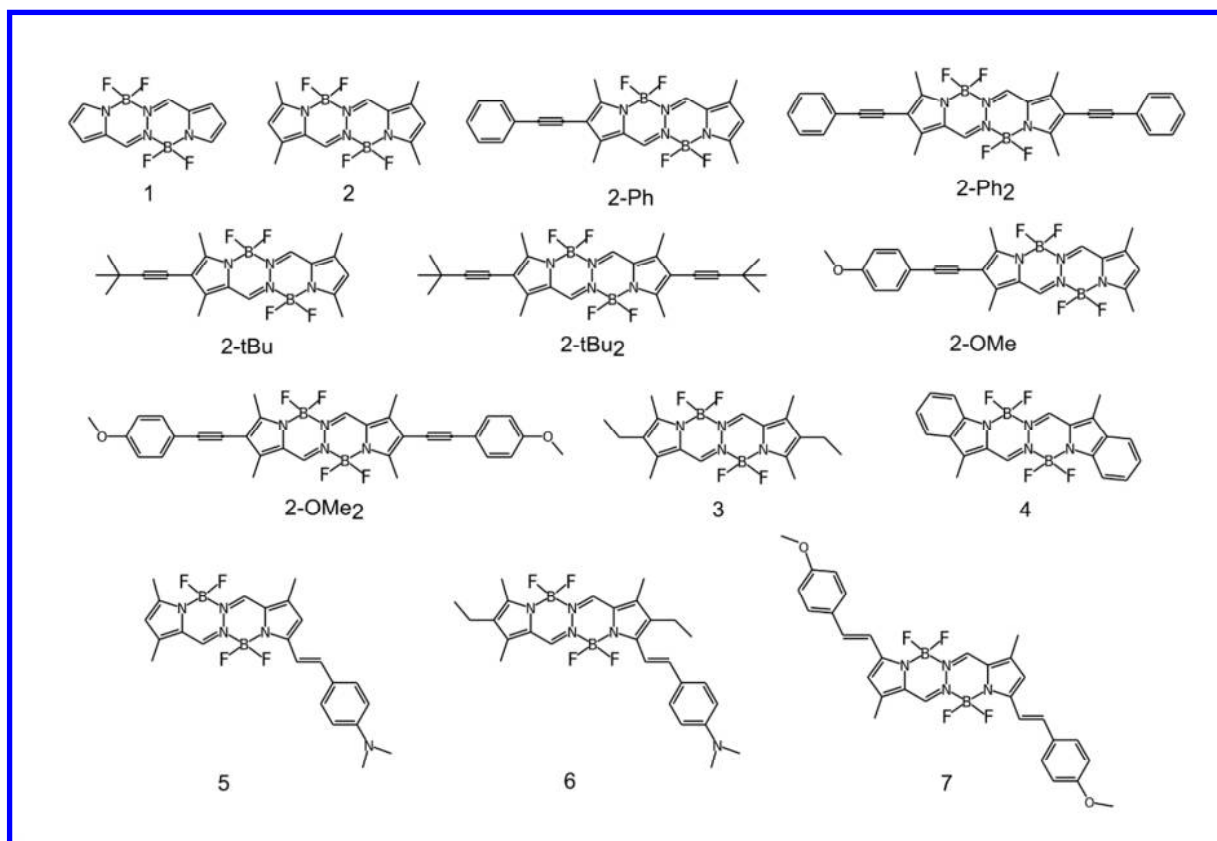
We have investigated a large number of BOPHY dyes that present exciting photophysical signatures for applications in several fields. The 0-0 energies trend for the studied BOPHY systems is well reproduced by a protocol relying on TD-DFT for excitation energies and cLR-PCM for solvent effects as evidenced by a R^2 of 0.968 when comparisons with experiments are made. However, this scheme yields large values of both the MAE (0.431 eV) and MSE (0.188 eV) compared to experiment. The subsequent incorporation of a SOS-CIS(D) correction of the transition energies considerably improves both the MAE (0.061 eV) and MSE (0.005 eV) values, at the cost of only a very slight deterioration of the determination coefficient (0.950). The cLR-TD-DFT/SOS-CIS(D) mixed approach therefore allows to reproduce all key experimental features with high accuracy and is a method of choice to model and design BOPHY derivatives.

A fine tuning of the spectroscopic properties of the BOPHY dyes can be achieved by structural modifications involving the incorporation of different substituents to the pyrrole units. Remarkably, the incorporation of one dimethylaminophenyl group via a vinyl linker in the α -position gives rise to two BOPHYs with particularly high absorption and emission maxima (**5** and **6**). We have rationalized these effects as resulting from a pronounced donor-acceptor behavior absent in the parent structures. These high absorption/emission maxima, along as the large Stokes shifts found for both systems, make them promising candidates for bioluminescence applications. It is therefore our hope that this work will stimulate the modeling of new BOPHY derivatives for biological applications.

FIGURES



Scheme 1. Typical Structure of a BOPHY dye.



Scheme 2. Representation of the complete set of BOPHY systems studied in the present work.

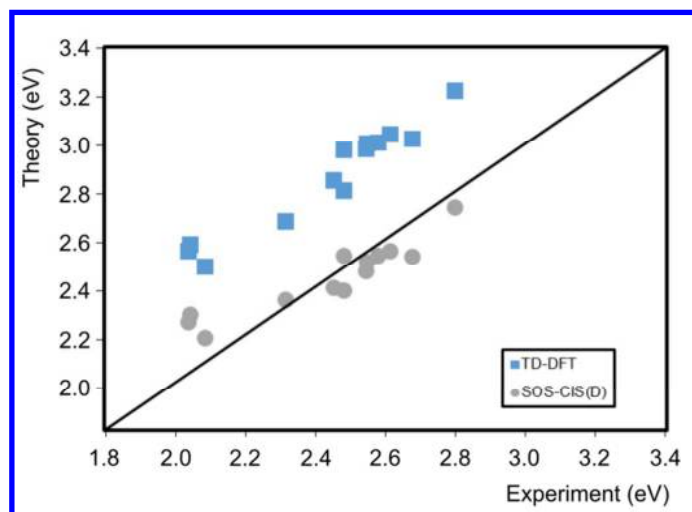


Figure 1. Comparison between theoretical values without considering solvent effects [TD-DFT and SOS-CIS(D)] and experimental AFCP values (eV). The central line indicates a perfect theory-experiment match.

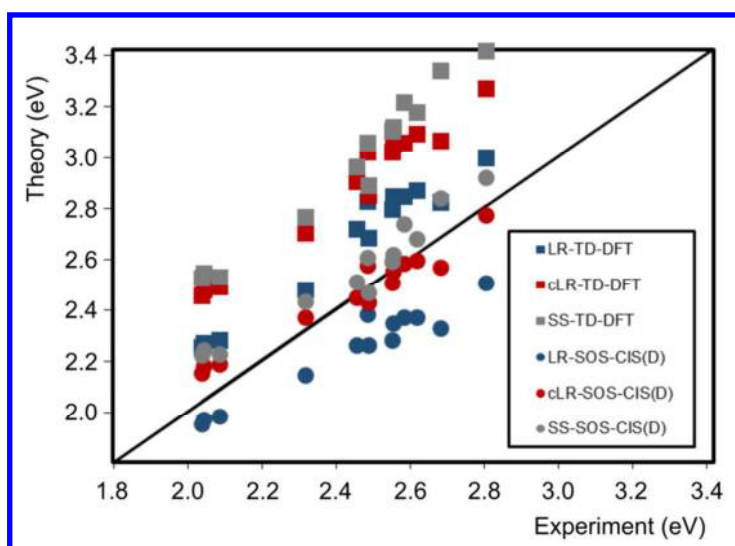


Figure 2. Comparison between theoretical (LR-TD-DFT, cLR-TD-DFT, SS-TD-DFT, LR-SOS-CIS(D), cLR-SOS-CIS(D) and SS-SOS-CIS(D)) and experimental AFCP values (eV). All calculations are carried out considering dichloromethane as solvent. The central line indicates a perfect theory-experiment match.

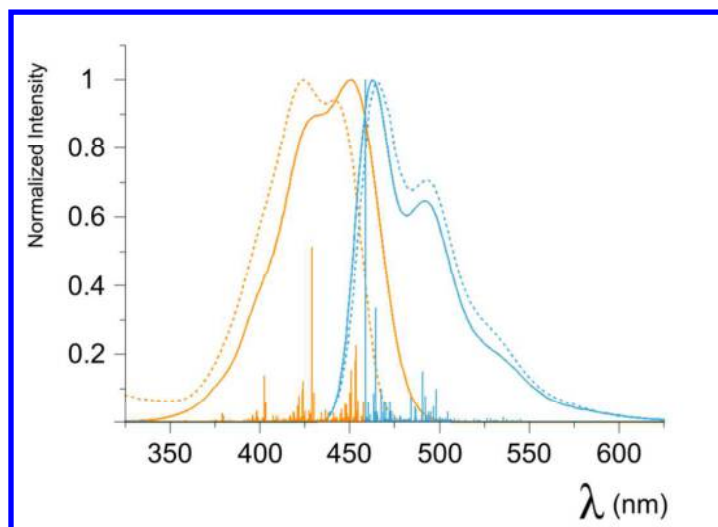


Figure 3. Comparisons between the experimental (dotted lined) and theoretical (full lines and sticks) normalized absorption (in orange) and emission (in blue) spectra of 1, in dichloromethane solution. The experimental curve is adapted from Ref. 11.

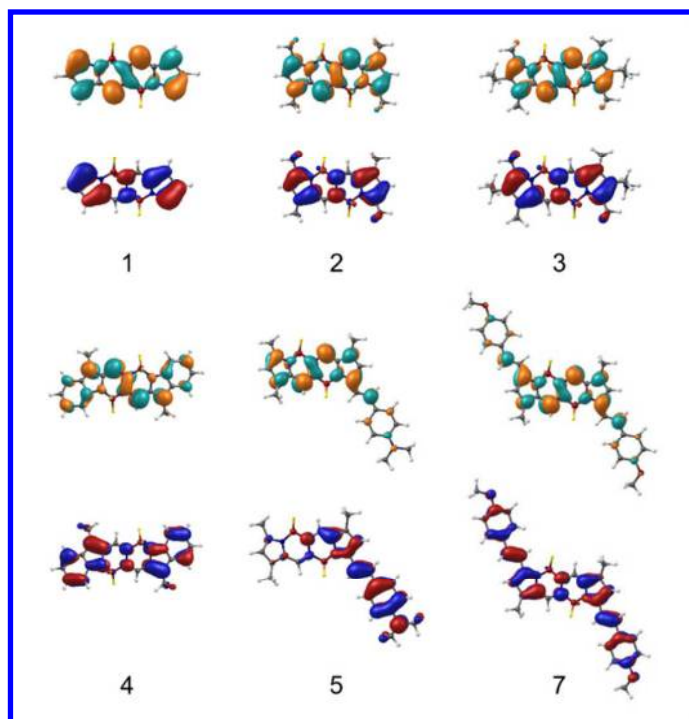


Figure 4. HOMO-LUMO plots for some selected chromophores. The HOMO is represented in red-blue, and the LUMO in turquoise-orange. A contour threshold of 0.03 a.u. has been applied.

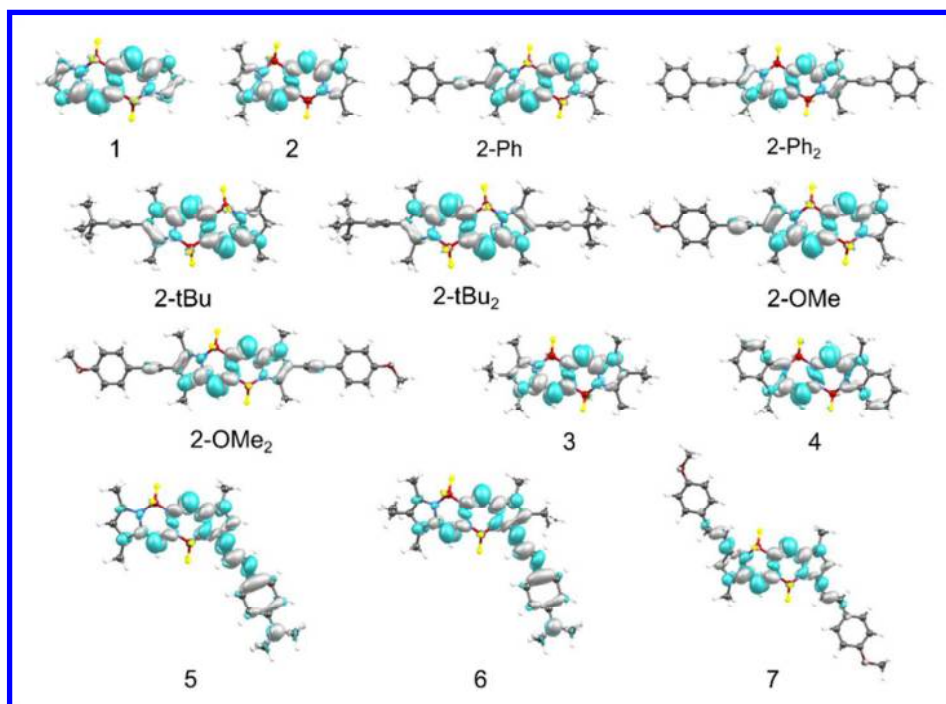


Figure 5. Electronic Density Difference plots for the studied systems. Turquoise and gray colors indicate density accumulation and depletion, respectively. A contour threshold of 0.001 a.u. has been applied.

TABLES

Table 1. Experimental AFCP energies (eV) and their corresponding calculated values using various computational models. All calculations were carried out considering dichloromethane as solvent.

	Exp ^a	Theory							
		TD-DFT	SOS-CIS(D)	PCM-TD-DFT			PCM-TD-DFT-SOS-CIS(D)		
				LR	cLR	SS	LR	cLR	SS
1	2.795	3.220	2.734	2.980	3.246	3.392	2.495	2.760	2.906
2	2.674	3.021	2.530	2.808	3.045	3.315	2.317	2.554	2.825
2-Ph	2.577	3.003	2.534	2.830	3.039	3.194	2.361	2.570	2.724
2-Ph₂	2.451	2.847	2.397	2.704	2.887	2.947	2.254	2.437	2.497
2-<i>t</i>Bu	2.611	3.041	2.551	2.852	3.072	3.156	2.362	2.582	2.666
2-<i>t</i>Bu₂	2.547	3.001	2.506	2.832	3.035	3.100	2.338	2.540	2.605
2-OMe	2.480	2.972	2.531	2.812	3.005	3.036	2.371	2.564	2.595
2-OMe₂	2.481	2.803	2.386	2.669	2.834	2.873	2.252	2.417	2.456
3	2.544	2.977	2.468	2.782	3.003	3.085	2.274	2.495	2.577
4	2.314	2.677	2.348	2.467	2.690	2.752	2.138	2.361	2.423
5	2.042	2.581	2.286	2.260	2.467	2.531	1.965	2.172	2.236
6	2.038	2.551	2.254	2.244	2.443	2.514	1.947	2.146	2.217
7	2.085	2.487	2.189	2.274	2.479	2.517	1.977	2.181	2.220

^a Experimental data from references 11, 12, 14 and 16

Table 2. Statistical analysis obtained from the comparison of experimental and theoretical AFCP energies (Table 1): mean absolute error (MAE), mean square error (MSE), standard deviation (SD), maximal positive and negative deviations [Max(+) and Max(-)], and theory/experiment linear determination coefficients (R^2). All values except those for R^2 are given in eV. All calculations except those concerning the first two lines [TD-DFT and SOS-CIS(D)] were carried out considering dichloromethane as solvent.

	MAE	MSE	SD	Max(+)	Max(-)	R^2
TD-DFT	0.426	0.185	0.224	0.538	0.322	0.935
SOS-CIS(D)	0.094	0.013	0.148	0.244	-0.144	0.869
LR-TD-DFT	0.221	0.052	0.255	0.332	0.133	0.955
cLR-TD-DFT	0.431	0.188	0.264	0.525	0.353	0.968
SS-TD-DFT	0.521	0.277	0.299	0.641	0.392	0.965
LR-SOS-CIS(D)	0.199	0.046	0.175	-0.077	-0.357	0.941
cLR-SOS-CIS(D)	0.061	0.005	0.185	0.130	-0.120	0.950
SS-SOS-CIS(D)	0.104	0.014	0.223	0.194	-0.025	0.936

Table 3. Selected ground-state and excited-state dihedral angles (degrees) and bond distances (Å) for the optimized structures of the studied BOPHYs

Dihedral	1	2	2-Ph	2-Ph₂	2-<i>t</i>Bu	2-<i>t</i>Bu₂	2-OMe
GS	29.1	28.3	28.6	28.9	28.6	28.7	28.7
ES	23.1	21.9	22.6	23.0	22.3	22.6	22.9
N-N							
GS	1.390	1.389	1.388	1.389	1.389	1.389	1.388
ES	1.326	1.328	1.329	1.330	1.328	1.328	1.331
N _(central) -B							
GS	1.602	1.596	1.593	1.317	1.317	1.318	1.593
ES	1.601	1.607	1.605	1.605	1.606	1.357	1.603
Dihedral	2-OMe₂	3	4	5	6	7	
GS	28.8	27.9	28.4	28.6	28.3	29.0	
ES	23.3	21.3	23.5	24.6	24.3	23.6	
N-N							
GS	1.388	1.388	1.387	1.387	1.386	1.385	
ES	1.331	1.328	1.336	1.350	1.349	1.337	
N _(central) -B							
GS	1.597	1.595	1.615	1.592	1.593	1.593	
ES	1.604	1.607	1.608	1.596	1.594	1.600	

Table 4. Experimental absorption maxima (nm), emission maxima (nm), Stokes shift (cm^{-1}), and calculated CT charge ($|e|$), distance (\AA), and dipole (Debye).

	Abs. max. (nm) Exp ^b .	Em. max. (nm) Exp ^b .	Stokes shift (cm^{-1}) Exp ^b .	CT charge ($ e $)	CT distance (\AA)	CT dipole (Debye)
1	424	465	2080	0.442	-	-
2	444	485	1904	0.431	-	-
2-Ph	456	509	2283	0.483	1.206	2.799
2-Ph₂	489	524	1366	0.518	-	-
2-<i>t</i>Bu	453	499	2035	0.465	0.723	1.614
2-<i>t</i>Bu₂	464	512	2020	0.490	-	-
2-OMe	479	523	1756	0.511	1.950	4.782
2-OMe₂	468	536	2711	0.540	-	-
3	478	497	800	0.446	-	-
4	498	580	2839	0.458	-	-
5	557	667	2961	0.603	3.563	10.317
6	557	670	3028	0.602	3.468	10.033
7	579	611	905	0.516	-	-

^b Experimental data from references 11, 12, 14 and 16

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7 Movies of the key vibrations (.avi files) are available free of charge at <http://pubs.acs.org>.
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